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Ceramic Nanofiltration Membrane for Use in Organic Solvents  
and Method for the Production Thereof

The invention concerns a ceramic nanofiltration membrane for use in organic solvents, as  
5 well as a method for its production.

Ceramic filter elements generally have an asymmetric structure in which thin membrane layers with one or more intermediate layers are applied to a porous ceramic support. Different membrane filtration ranges, like microfiltration, ultrafiltration and  
10 nanofiltration are distinguished according to the pore size or retention capacity.

The porous ceramic support stipulates the external shape and mechanical stability of the filter element. Common variants are disks, which are produced by film casting or pressing, and tubes, which are extruded in most cases as rigid plastics. After the sintering  
15 process, which can lie between 1200 and 1700°C, depending on the employed ceramic material, an open-pore ceramic body is obtained. The pores are formed by cavities between the sintered grains (intermediate grain pores). Pore sizes between 1 µm and 10 µm can be set, depending on the employed initial particle size and particle shape.

20 Ceramic microfiltration membranes in most cases are produced using narrowly classified ceramic powders. These powders are dispersed in an appropriate solvent, using auxiliaries (dispersants). The slurry so produced is mixed with solutions of organic binders and then used to coat the porous ceramic support. In contact with the porous support the solvent penetrates into the support, forming a thin ceramic layer on the  
25 surface. This is dried and then fired at high temperature between 800 and 1400°C, depending on the employed fineness of the powder (**A. Burggraaf, K. Keizer, in Inorganic Membranes, ed. R. R. Bhave, Van Nostrand Reinhold, New York, p. 10-63**). Increasingly more fine-pored membranes can be produced with increasingly finer ceramic powders in this way. The finest available powders have a particle size of  
30 about 60-100 nm, from which membranes with a pore size of about 30 nm can be produced. This is the upper range of ultrafiltration.

Ceramic ultrafiltration membranes with a lower separation limit can be produced via the sol-gel technique in aqueous solution. For this purpose, metal alcoholates are preferably  
35 fully hydrolyzed in water. Colloidal hydroxide and hydrated oxide particles are formed, which can be stabilized by adding small amounts of electrolyte (mineral acid or alkali). Solutions of organic binders are added to these so-called sols and the solution so formed used for coating. Ceramic microfiltration membranes or coarse ultrafiltration membranes

are used as supports. The solvent of the sol penetrates into the porous support. The concentration increase on the surface leads to a sharp rise in viscosity and formation of a gel. This gel is then carefully dried and finally fired at temperatures between 400 and 1000°C. In this way ceramic membranes can be produced for the lower range of ultrafiltration with slit-like pores and a pore diameter between 3 nm and 10 nm (**I. Voigt, G. Mitreuter, M. Füting, CfI/Ber. DKG 79 (2002), E39-E44**).

Ceramic membranes for nanofiltration can be produced via a special form of the sol-gel technique. For this purpose an organic solvent is used instead of water and hydrolysis is carried out only with a defined amount of water, which is substoichiometric relative to the number of hydrolyzable alcoholate groups. At the same time, the formed hydroxide groups begin to condense with water cleavage. The oligomers so produced are inhibited from further chain growth with increasing chain length and the state of a sol is also formed, which is also referred to in this special case as a polymer sol (**C. J. Brinker, G. W. Scherer, Sol-Gel Science, Academic Press, Inc., 1990**). Because of strong dilution of the sol this state can be stabilized for several days. Addition of binders is not required owing to the polymer structure of the sol. The principle of layer formation is again comparable to that in ceramic slurries and particulate salts. The solution penetrates into the pores, the viscosity of the surface layer increases, a gel is formed. This is dried and then fired at temperatures between 200 and 600°C. In this way cylindrical pores with an average pore diameter of less than 2 nm are obtained. Voight et al. (**Proc. of 5<sup>th</sup> Inter. Conf. on Inorg. Membr. (5<sup>th</sup> ICIM), July 22-26, 1998, Nagoya, Japan, pp. 42-45**) produced in this way TiO<sub>2</sub> nanofiltration membranes with an average pore diameter of 0.9 nm, a water flow rate of 20 L/(m<sup>2</sup>·h·bar) and a molecular separation limit (90% retention) in aqueous solution at 450 g/mol (**J. Membr. Sci. 174 (2000), 123-133**).

**WO98/17378** describes an inorganic nanofiltration membrane consisting of sintered metal oxide particles in a graded layer sequence on a monolithic, ceramic, multichannel support. It carries on the channel walls a microfiltration layer, this has an ultrafiltration layer and this finally a nanofiltration layer, whose equivalent pore diameter before sintering lies in the range between 0.5 nm and 1.5 nm and has a separation edge between 100 and 2000 dalton. The nanofiltration membrane preferably consists of zirconium oxide and is preferably produced in the sol-gel method by hydrolysis in an alcoholic media. The area of application is processing of salt solutions (for example NaCl solutions) during regeneration of ion exchange resins, which are used in refining of sugar cane. Aqueous solutions are therefore involved here, not organic solvents.

Depending on the pore size, a distinction is made between macropores with a pore size >100 nm, mesopores with a pore size between 100 nm and 2 nm and micropores with a pore size less than 2 nm. Ceramic ultrafiltration membranes therefore preferably have mesopores, ceramic nanofiltration membranes and micropores.

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If one investigates the described nanofiltration membranes with respect to their separation behavior in organic solvents, one surprisingly finds that the flow rate in comparison with water drops sharply. A flow rate  $<5 \text{ L}/(\text{m}^2 \cdot \text{h} \cdot \text{bar})$  is measured in the TiO<sub>2</sub> membranes with a pore size of 0.9 nm. Separation limit increases at the same time to  $>2000 \text{ g/mol}$ .

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In order to obtain ceramic nanofiltration membranes for use in organic solvents, an attempt was made to improve the flow rate of the organic solvent through the membrane pores, using mixed oxides.

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Guizard et al. (**Desalination 147 (2002), 275-280**) investigated the mixed oxides SiO<sub>2</sub>/TiO<sub>2</sub>, Al<sub>2</sub>O<sub>3</sub>/ZrO<sub>2</sub> and SiO<sub>2</sub>/ZrO<sub>2</sub>. They obtained microporous ceramic membranes with pore radii  $\leq 1 \text{ nm}$  with scarcely improved permeation behavior.

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Tsuru et al. (**J. Coll. Interf. Sci. 228 (2000), 292-296; J. Membr. Sci. 185 (2001), 253-261**) investigated the behavior of SiO<sub>2</sub>/ZrO<sub>2</sub> membranes produced via a sol-gel process. They varied the pore size between 1 nm and 5 nm. This did not lead to a flow rate as obtained in aqueous solvents either.

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Our own studies showed that the reason for this behavior lies in the strong hydrophilicity of the ceramic micropores, which is caused by the fact that water or OH groups are added to the oxide surface. These micropores are not permeable to organic solvent molecules. Transport occurs through larger pores and/or defects, which only have a limited fraction of the total pore volume. Because of this the flow rate drops in comparison with water flow. The retention of these larger pores or defects lies well above that of the micropores.

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If the micropores are configured hydrophobic, permeation of water should be inhibited and possibly completely restricted. The flow rate of organic solvents is surprisingly improved and the retention of molecules dissolved in organic solvents is not improved relative to hydrophilic NF membrane.

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**WO92/06775** claims a nanofiltration membrane in which a support consisting of an inorganic substance is coated with a first layer with a porous inorganic material with a

pore radius lower than 10 nm and contains a second active layer with a thickness of 0.1 µm to 1 µm, which consists of an organic polymer.

5       **WO99/61140** describes a method for production of hydrophobic inorganic membranes by a sol-gel process, in which alcoholates with at least one nonhydrolyzable hydrocarbon group are used. These are added to the hydrolysis process after hydrolysis of the pure alcoholates has progressed to a certain point. As an example of alcoholates with nonhydrolyzable hydrocarbon groups, methyltriethoxysilane is described. A sol is formed that contains the hydrophobing components. As a result, microporous hydrophobic  
10      membranes with pore sizes of 0.5 nm and 0.7 nm are obtained, which can be used for gas separation.

15       **WO99/29402** describes an inorganic filter membrane consisting of a support coated with a membrane, which contains covalently bonded organomineral or mineral titanium or zirconium groups.

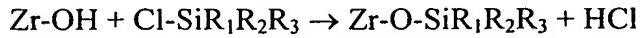
The underlying task of the invention is to devise a ceramic nanofiltration membrane for use in organic solvents, as well as a method for its production, while avoiding the deficiencies and complicated procedures of the prior art.

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This task is solved by the invention described in the claims.

In contrast to the prior art just outlined, according to the present invention the pores of mesoporous ceramic membranes, which are ordinarily used for ultrafiltration, are modified by subsequent treatment with a hydrophobing agent. The pore size then lies between 2 nm and 10 nm, preferably between 2 nm and 5 nm. The membrane consists of oxides of aluminum, silicon, zirconium or titanium or their mixtures. A silane with the general formula R<sub>1</sub>R<sub>2</sub>R<sub>3</sub>R<sub>4</sub>Si is preferably used as hydrophobing agent, in which between 25      one and three, preferably one of the groups R<sub>1</sub>-R<sub>4</sub> are hydrolyzable groups, like -Cl, -OCH<sub>3</sub> or -O-CH<sub>2</sub>-CH<sub>3</sub>. The remaining groups are nonhydrolyzable groups like alkyl groups, phenyl groups, which can be at least partially fluorinated to increase the hydrophobic effect. Bonding of the hydrophobing agents to the membrane surface occurs 30      by a condensation reaction of the hydrolyzable groups with OH groups on the surface of the oxide membrane according to the following reaction equation:

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The use of a hydrophobing agent with only one hydrolyzable substituent preferably leads to a situation in which a monomolecular layer is added without the molecules of a hydrophobing agent reacting with each other.

- 5 Modification of the mesoporous ceramic membranes using the described hydrophobing agent can occur either in the liquid phase by impregnation of the membrane in a solution of the hydrophobing agent or in the gas phase by applying a vacuum or using a carrier gas.
- 10 By applying a pressure difference between the front and back side of the membrane, penetration of the hydrophobing agent can be supported. For better fixation of the hydrophobing agent, heat treatment between 100 and 400°C, preferably between 150 and 300°C can be used at the end.
- 15 The invention is further explained below on three practical examples. The accompanying drawings show:

Figure 1: Permeate flow of a 3 nm ZrO<sub>2</sub> membrane without and with hydrophobing with n-octyldimethylchlorosilane in the liquid phase (example 1),

20 Figure 2: Permeate flow of a 5 nm ZrO<sub>2</sub> membrane without and with hydrophobing with tridecafluoro-1,1,2,2-tetrahydrooctyl-trichlorosilane in the liquid phase with and without vacuum support (example 2) and

25 Figure 3: Permeate flow of a 3 nm ZrO<sub>2</sub> membrane without and with hydrophobing with trimethylchlorosilane in the gas phase (example 3).

30 Example 1: Hydrophobing with n-octyldimethylchlorosilane in the liquid phase

n-Octyldimethylchlorosilane is a molecule with a hydrolyzable group (-Cl), two nonhydrolyzable methyl groups (-CH<sub>3</sub>) and a nonhydrolyzable octyl group (-(CH<sub>2</sub>)<sub>7</sub>-CH<sub>3</sub>).

35 1 g of this compound is dissolved in 100 g n-heptane.

A mesoporous ZrO<sub>2</sub> membrane (manufacturer inocermic GmbH) with an average pore size of 3 nm is immersed in this solution. After a residence time of 2 minutes, the

membrane is removed from the solution, dried in air for 10 minutes and then treated in a drying cabinet for 30 minutes at 175°C.

Hydrophobing shows up through a different wetting behavior relative to water. The  
5 untreated ZrO<sub>2</sub> membrane is wetted so well that a contact angle <10° is measured. After hydrophobing the surface has a contact angle of 80°.

Investigation of solid flow occurs in the crossflow method at a transmembrane pressure  
of 3 bar and an overflow rate of 2 m/s. Hydrophobing with n-octyldimethylchlorosilane  
10 leads to a sharp reduction in water flow from 95 L/(m<sup>2</sup>·h·bar) to 1.5 L/(m<sup>2</sup>·h·bar). In contrast to this, methanol flow rises from 32 L/(m<sup>2</sup>·h·bar) to 51 L/(m<sup>2</sup>·h·bar) and toluene flow from 18 L/(m<sup>2</sup>·h·bar) to 22 L/(m<sup>2</sup>·h·bar).

Separation limit measurements with polystyrene standards in toluene at a transmembrane  
15 pressure of 3 bar and an overflow rate of 2 m/s give a separation limit (90% retention) of 1025 g/mol for the nonhydrophobized 3 nm ZrO<sub>2</sub> membrane. After hydrophobing, the separation limit drops to 660 g/mol.

20 Example 2: Hydrophobing with tridecafluoro-1,1,2,2-tetrahydrooctyltrichlorosilane in the liquid phase

Tridecafluoro-1,1,2,2-tetrahydrooctyltrichlorosilane is a molecule with three hydrolyzable groups (-Cl) and a long-chain, strongly fluorinated nonhydrolyzable group. 1 g of this  
25 compound is dissolved in a mixture of 50 g ethanol (99.8%) and 50 g heptane.

A mesoporous TiO<sub>2</sub> membrane with an average pore size of 5 nm (manufacturer inocermic GmbH) is immersed in this solution. After a residence time of 2 minutes the membrane is removed from the solution, dried in air for 10 minutes and then treated in a  
30 drying cabinet for 30 minutes at 175°C.

Hydrophobing shows up by a different wetting behavior relative to water. The untreated TiO<sub>2</sub> membrane is wetted so well that a contact angle <10° is measured. After hydrophobing, the surface has a contact angle of 120°.

35 The investigation of solvent flow occurs in the crossflow method at a transmembrane pressure of 3 bar and an overflow rate of 2 m/s. Hydrophobing with tridecafluoro-1,1,2,2-tetrahydrooctyltrichlorosilane leads to a sharp reduction of water flow from 90

$L/(m^2 \cdot h \cdot bar)$  to  $2 L/(m^2 \cdot h \cdot bar)$ . In contrast to this the methanol flow increases from  $40 L/(m^2 \cdot h \cdot bar)$  to  $75 L/(m^2 \cdot h \cdot bar)$  and toluene flow from  $10 L/(m^2 \cdot h \cdot bar)$  to  $30 L/(m^2 \cdot h \cdot bar)$ .

- 5 Separation limit measurements of polystyrene standards in toluene at a transmembrane pressure of 3 bar and an overflow rate of 2 m/s give a separation limit (90% retention) of 1800 g/mol for the nonhydrophobized 5 nm  $TiO_2$  membrane. After hydrophobing, the separation limit drops to 1200 g/mol.
- 10 Hydrophobing can be supported by applying a vacuum to the membrane back side and drying the hydrophobing agent into the membrane pores. If a vacuum of 60 mbar is applied during hydrophobing with 1% n-octyldimethylchlorosilane solution in n-heptane, no water flow can be measured under the described experimental conditions. The methanol flow is  $50 L/(m^2 \cdot h \cdot bar)$ , toluene flow  $20 L/(m^2 \cdot h \cdot bar)$ .

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Example 3: Hydrophobing with trimethylchlorosilane in the gas phase

Trimethylchlorosilane has one hydrolyzable group (-Cl) and three methyl groups that are used for hydrophobing.

A glass dish with trimethylchlorosilane is placed on the bottom of a closable container. The weighed amount of trimethylchlorosilane is 2 g per square decimeter of surface being coated. Mesoporous  $ZrO_2$  membranes (manufacturer inocermic GmbH) with a pore size of 3 nm are arranged in the container above the hydrophobing agent so that no contact with trimethylchlorosilane exists. The container is closed and evacuated with a membrane pump to a pressure of about 250 mbar, during which trimethylchlorosilane begins to boil at room temperature. The pump is switched off. The pressure in the container rises from the evaporating trimethylchlorosilane. After a waiting time of 10 minutes, it is evacuated again. This procedure is repeated a total of three times (at room temperature). The membranes are then tempered in air for 1 hour at  $150^\circ C$ .

Hydrophobing shows up by a different wetting behavior relative to water. The untreated  $ZrO_2$  membrane is wetted so well that a contact angle of  $<10^\circ$  is measured. After hydrophobing, the surface has a contact of  $40^\circ$ .

Investigation of solvent flow occurs in the crossflow method at a transmembrane pressure of 3 bar and an overflow rate of 2 m/s. Hydrophobing with trimethylchlorosilane leads to

a sharp reduction of water flow from 95 L/(m<sup>2</sup>·h·bar) to 4.5 L/(m<sup>2</sup>·h·bar). In contrast to this methanol flow rises from 32 L/(m<sup>2</sup>·h·bar) to 45 L/(m<sup>2</sup>·h·bar) and toluene flow from 18 L/(m<sup>2</sup>·h·bar) to 30 L/(m<sup>2</sup>·h·bar).

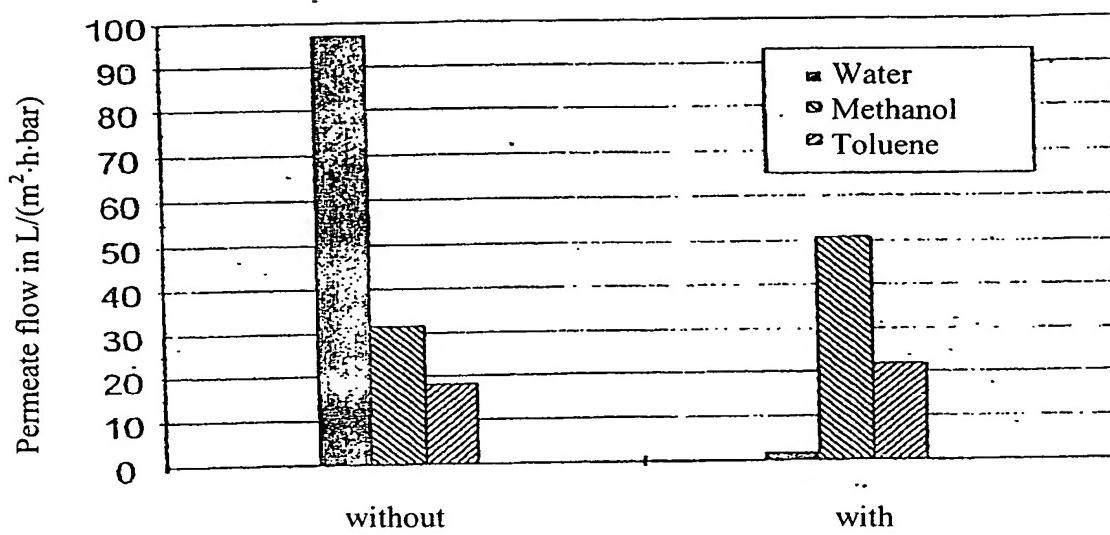
- 5 Separation limit measurements for polystyrene standards in toluene at a transmembrane pressure of 3 bar and an overflow rate of 2 m/s gives a separation limit (90% retention) of 1025 g/mol for the unhydrophobized 3 nm ZrO<sub>2</sub> membrane. After hydrophobing, the separation limit drops to 800 g/mol.

## Claims

1. Ceramic nanofiltration membrane for use in organic solvents, characterized by the fact that a mesoporous ceramic membrane ordinarily used for ultrafiltration is modified by treatment with a hydrophobing agent.
2. Ceramic membrane according to Claim 1, characterized by the fact that the pore size of the mesoporous membrane is between 2 nm and 10 nm, preferably between 2 nm and 5 nm.
3. Ceramic membrane according to one of the preceding claims, characterized by the fact that the mesoporous ceramic membrane consists of a metal oxide, preferably  $TiO_2$ ,  $ZrO_2$ ,  $Al_2O_3$  or  $SiO_2$  or mixtures of two or more of these oxides.
4. Ceramic membrane according to one of the preceding claims, characterized by the fact that the hydrophobing agent used for modification is a silane of the general formula  $R_1R_2R_3R_4Si$ .
5. Ceramic membrane according to Claim 4, characterized by the fact that between one and three, but preferably one of the groups  $R_1-R_4$  are hydrolyzable groups, like -Cl, -OCH<sub>3</sub> or -O-CH<sub>2</sub>-CH<sub>3</sub>.
6. Ceramic membrane according to Claim 4, characterized by the fact that between one and three but preferably three of the groups  $R_1-R_4$  are nonhydrolyzable groups, like alkyl groups, phenyl groups.
7. Ceramic membrane according to Claim 6, characterized by the fact that, to increase the hydrophobic effect, at least one of the nonhydrolyzable substituents is at last partially fluorinated.
8. Method for production of a ceramic membrane according to one of the preceding claims, characterized by the fact that modification of the mesoporous membrane occurs by impregnation with the hydrophobing agent in the liquid phase.
9. Method according to Claim 8, characterized by the fact that penetration of a hydrophobing agent is supported by a pressure difference between the front and back side of the membrane.
10. Method for production of a ceramic membrane according to one of the Claims 1 to 7, characterized by the fact that modification of the mesoporous membrane occurs with the hydrophobing agent from the gas phase.

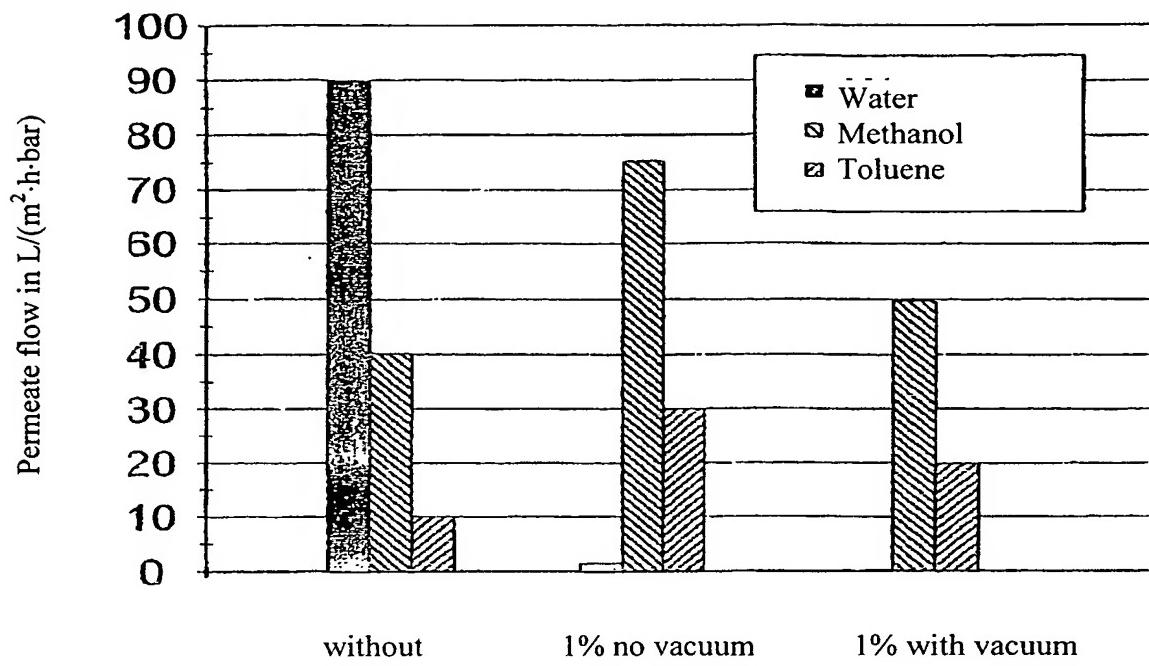
11. Method according to one of the Claims 1 to 9, characterized by the fact that, after treatment with the hydrophobing agent, heat treatment between 100 and 400°C, preferably between 150 and 300°C is applied.

Figure 1.



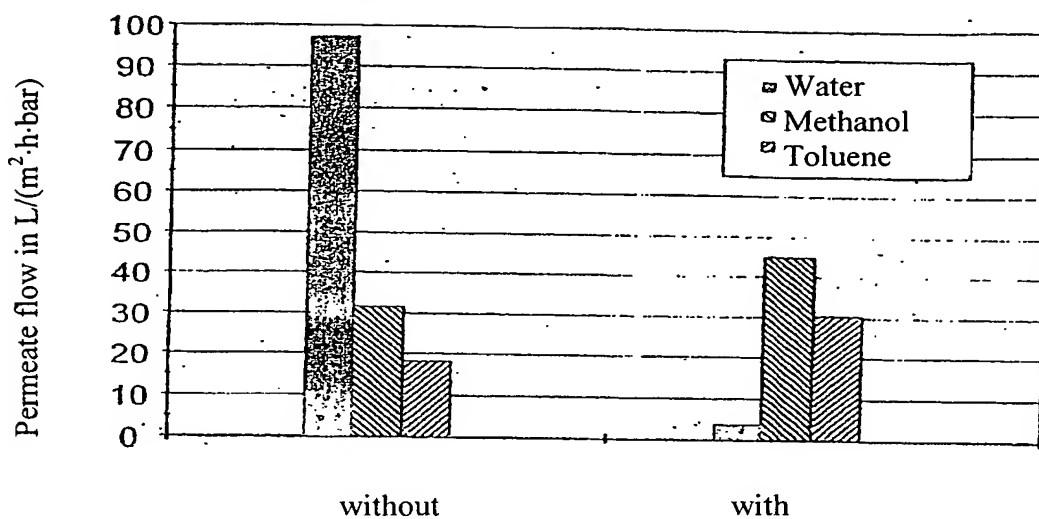
Permeate flow of a 3 nm  $\text{ZrO}_2$  membrane without and with hydrophobing with n-octyldimethylchlorosilane in the liquid phase.

Figure 2.



Permeate flow of a 5 nm TiO<sub>2</sub> membrane without hydrophobing and with hydrophobing with tetradecafluoro-1,1,2,2-tetrahydrooctyltrichlorosilane in the liquid phase with and without vacuum support.

Figure 3.



Permeate flow of a 3 nm  $\text{ZrO}_2$  membrane without and with hydrophobing with trimethylchlorosilane in the gas phase.